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(not in text) Key

Concepts: Terms in this set (13) Chemical Equilibrium. A state in which the rates of forward and reverse reactions are the same. An active, dynamic condition, not static.

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ANS: The reversible reaction always attains equilibrium which proceeds both sides and never goes for completion. 2. The reaction  $\text{CaCO}_3 \rightleftharpoons \text{CaO} + \text{CO}_2(\text{g})$  goes to completion in lime because: a) Of the high temperature. b)  $\text{CaO}$  is more stable than  $\text{CaCO}_3$ . 3. c)  $\text{CaO}$  is not

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Chemical Equilibrium -  
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View Answer. Topic:

UNIT 8: Physical and  
Chemical Equilibrium

Tag: TN 11th

Chemistry. Q. 8. In the  
equilibrium,  $2A(g)$

$2B(g) + C_2(g)$  the  
equilibrium

concentrations of A, B  
and  $C_2$  at 400 K are 1



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$\times 10^{-4} \text{ M}$ ,  $2.0 \times 10^{-3} \text{ M}$ ,  $1.5 \times 10^{-4} \text{ M}$  respectively. The value of  $K_c$  for the equilibrium at 400 K is.

Topic: UNIT 8:

Physical and Chemical  
Equilibrium (Test 1)

General form of an  
equilibrium reaction:  $aA + bB \rightleftharpoons cC + dD$

“ A & B ” are  
reactants, “ C & D ”

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are products, and “ a, b, c, d ” are coefficients.

To determine the equilibrium constant (K) from the general model:

$$K = \frac{[C]^c [D]^d}{[A]^a [B]^b}$$

If a K-value greater than 1 is obtained, more products than reactants at equilibrium. If a K-value less than one is obtained, there are more reactants than

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products at equilibrium.

## Chemical Equilibrium Test Answers

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If all are equal, answer

E. a. the

concentrations of

reactant and products

are equal b. the rate

constants for the forward

and reverse reactions

are equal c. the time that

a particular atom or

molecule spends as a

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Chemical Equilibrium  
Test Answers

reactant and product  
are equal d. the rate of  
the forward and reverse  
reaction e. all of the  
above are equal

Big-Picture Introductory  
Conceptual Questions

d) equilibrium always  
shifts to form fewer  
molecules Answers : 1)  
d, 2) c, 3) c, 4) c, 5) a, 6)  
a, 7) c, 8) d, 9) b, 10) c,  
11) d, 12) b, 13) b, 14) d,

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15) b, 16) b, 17) A  
chemical equilibrium  
exists when the forward  
rate equals the reverse  
rate for a reversible  
chemical reaction., 18)  
b, 19) b, 20) b, 21) d.

Chem12 Equilibrium :  
Exam Questions 1-50

18) The following  
equilibrium is readily  
established:  $\text{SO}_2\text{Cl}_2 (\text{g})$   
 $\rightleftharpoons \text{SO}_2 (\text{g}) + \text{Cl}_2 (\text{g})$  At

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equilibrium at 373 K, a  
1.00-L reaction vessel  
contains 0.0106 mol of  
 $\text{SO}_2\text{Cl}_2$  and 0.0287 mol  
each of  $\text{SO}_2$  and  $\text{Cl}_2$ .

What is  $K_{\text{eq}}$  for the  
reaction at 373 K?

- A)12.8 B)2.72 C)0.0781  
D)2.39 E)0.418

A.P. Chemistry Practice  
Test - Ch. 13:  
Equilibrium ...  
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Equilibrium Mcqs

Online Test 5

Preparation Question  
Answer. If a buffer

solution of higher pH  
than seven is to be made  
we use? Strong acid and  
strong base. Weak acid  
and strong base. Weak  
acid and strong base.

Chemistry Chemical  
Equilibrium Mcqs  
Online Test 5 ...

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Unit #3: Chemical



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Systems and

equilibrium. Thursday,  
November 7, 2019

Equilibrium Lab:

Equilibrium Answer  
Questions Practice Q

#1-6 pg. 422. Friday,  
November 8, 2019

Equilibrium Constants

PP Q#1-10 pg. 428,

Q#11-15 pg. 430,

Q#31-40 pg. 444

Answers. ... Unit Test

After Test Activity 1.2

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pg. 23 Sniffing out  
Cancer - not Q4 ...

Equilibrium  
Test Answers

Unit 3: Chemical  
Systems and  
Equilibrium - MS.

SWARTZ

2. A chemical reaction  
 $A \rightleftharpoons B$  is said to be in  
equilibrium when: Rate  
of transformation of A  
to B is equal to B to A.  
Complete conversion of  
A to B has taken place.

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Conversion of A to B is 50% complete. 50% reactant has been changed to B. Question 2 of 15.

Chemistry Chemical Equilibrium Online Quiz Test MCQs  
For a reaction involving only gases at 25°C the equilibrium constant can be expressed in terms of molarity  $K_c$  or

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partial pressure  $K_p$ .

Which is true about the numerical value of  $K_p$ ?

$K_c$  is generally equal to  $K_p$   
 $K_c$  is equal to  $K_p$  if the total moles of reactants and products are equal

PTS Chemistry  
Chemical Equilibrium  
Mcqs Online Test 5 ...  
Answers appear at the end of the test. Question

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1 An equilibrium constant with a value  $K > 1$  means:

- a. there are more reactants than products at equilibrium
- b. there are more products than reactants at equilibrium
- c. there are the same amount of products and reactants at equilibrium
- d. the reaction is not at equilibrium

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Equilibrium Constants

Practice Test -

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In a chemical

equilibrium, the rate constant for the forward reaction is  $2.5 \times 10^2$  and the equilibrium constant is 50. The rate constant for the reverse reaction is, a) 11.5 b) 5

Choose the best answer:

Chemistry: Physical and

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The concept of equilibrium permeates nearly every reaction that we will study in this course. Many reactions seek to achieve equilibrium and this unit represents the first look that we will take at how to characterize aqueous and gaseous equilibrium systems through a wide variety

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of chemical reactions.

Equilibrium

Unit 2: Equilibrium -

Ms. Bunney's Classes

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chemistry - Chemical



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## Equilibrium

### Chemical Equilibrium

shishir 2018-11-19T12:3  
5:50+05:30. This

section contains 15 questions(1-15). Each question carries 4 marks. For every incorrect answer 1 mark would be deducted.

Each question is having only **ONE CORRECT** answer .In order to secure any mark, correct

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... Take Test Now. This  
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in 15 ...

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